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# **Chemfiesta Ph Practice Worksheet Answers**

**Chemfiesta Ph Practice Worksheet  
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pH: 2.9 pOH: 11.1 2) What is the pH and pOH of a  $2.34 \times 10^{-5}$  NaOH solution?

pOH: 4.6 pH: 9.4 3) What is the pH and pOH of a solution made by adding water to 15 grams of hydroiodic acid until the volume of the solution is 2500 mL? pH: 1.6 pOH: 12.4

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## **School District**

To find the answer, take the negative log of this to find that the pH = 0.34 2) pH = 1.55 3) pH = 2.53 4) The pH of this solution is 6.35, making the solution very slightly acidic. 5) The pH will be 6. This is solved in the same way that dilution problems are solved. If the pH = 4, this means that the concentration of

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[H+] present is 0.0001 M.

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takes 25 mL of 0.05 M HCl to neutralize 345 mL of NaOH solution, what is the concentration of the NaOH ...

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Calculations Find the pH of the following  
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this answer is the limiting reagent  
Solutions to the Molarity Practice  
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## **Molarity Practice Answer Key Chemfiesta**

Mole Conversions and Stoichiometry  
Review Worksheet. 1) Using the following  
equation:  $2 \text{ NaOH} + \text{H}_2 \text{ SO}_4 \rightarrow 2 \text{ H}_2 \text{ O} +$

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Na<sub>2</sub>SO<sub>4</sub>. 4. How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid (H<sub>2</sub>SO<sub>4</sub>)?. 2) Using the following equation:  
$$\text{Pb}(\text{SO}_4)_2 + 4 \text{LiNO}_3 \rightarrow \text{Pb}(\text{NO}_3)_4 + 2 \text{Li}_2\text{SO}_4$$
4. How many grams of lithium nitrate will ...

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