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chemical reaction using the relationships found in the balanced chemical equation.

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Chemistry: Supplemental Stoichiometry Problems.

Directions: Solve each of the following problems.

Assume excess of any reactant that isn't mentioned, unless otherwise . specified. Show your work, including proper units, to earn full credit. 1. Given the balanced equation, show what the following molar ratios (i.e., the conversion factors) should be.

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Stoichiometry Packet Answer Key Answer: 4.93×10^{-5} L or 49.3 μ L In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. Stoichiometry Packet Answers Chapter 12 Chapter 12 Stoichiometry Packet Answers

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